**Name…………………………………………………….Index No…………………./…………**

**School…………………………………………Adm No…………………Stream………………**

**Date………………………………………………………Sign…………………………………..**

233/3

**CHEMISTRY**

Paper 2

(THEORY)

NOVEMBER/DECEMBER 2021

**TIME: 2 HOURS**

**SAMIA SUB-COUNTY JOINT EXAMINATION-2021**

***Kenya Certificate of Secondary Education (K.C.S.E) Trial Examination***

**CHEMISTRY**

**PAPER 2**

**INSTRUCTIONS TO CANDIDATES**

* *Write your* ***name****,* ***School*** *and* ***Index Number*** *in the spaces provided above.*
* ***Sign*** *and* ***write date*** *of examination in the spaces provided above.*
* *Answer* ***ALL*** *questions in the spaces provided.*
* *Mathematical tables and silent non-programmable electronic calculators may be used.*
* *All working* ***MUST*** *be clearly shown where necessary.*
* *Candidates should answer the questions in* ***English***
* *Candidates should check the question paper to ascertain that all pages are printed as indicated and that no questions are missing*
* *This paper contains 9 printed pages*

**For Examiner’s Use Only**

|  |  |  |
| --- | --- | --- |
| **Questions**  | **Maximum score** | **Candidate’s Score** |
| **1** | **12** |  |
| **2** | **10** |  |
| **3** | **13** |  |
| **4** | **11** |  |
| **5** | **11** |  |
| **6** | **11** |  |
| **7** | **12** |  |
| **Total** | **80** |  |

1. The information below relates to element N, P,Q, Rand S . Study it and answer the questions that follow. The letters are not the actual symbols for the elements.

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| Element | Atomic radius(mm) | Ionic radius(mm) | Formula of oxide | Melting point of oxide |
| N | 0.364 | 0.421 | N2O | -119 |
| P | 0.830 | 0.711 | PO2 | 837 |
| Q | 0.592 | 0.485 | Q2O3 | 1466 |
| R | 0.381 | 0.446 | R2O5 | 242 |
| S | 0.762 | 0.676 | SO | 1054 |

1. Name the elements that are metal. **Give** a reason. (2mks)
2. Compare the melting points of the oxides of S and R in terms of structure and bonding. (2mks)
3. Name the pair of elements that would react most vigorously with each other?

**Explain**  (2mks)

1. The table bedlow has information about chlorides of elements in period 3 of the periodic table:

Sulphur to sulphur

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| **Chloride**  | NaCI | MgCI2 | AICI3 | SiCI4 | PCI5 |
| **Melting point (oC)** | 801 | 712 | Sublimes at 183 | -70 | -80 |

What are the possible PH values of the solutions formed when the following chlorides are dissolved in water? Explain

MgCI2 (1mk)

AICI3 (1mk)

1. The molecular formula of Aluminum chloride is AI2CI6. Draw the structural (not dot and cross diagram) of Aluminum chloride indicating clearly the different types of bonds present. (2mks)
2. Using dot ( ) and cross (×), draw a diagram to show bonding in sodium chloride. (Na=11, CI=17) (2mks)
3.
4. What is the molar enthalpy of neutralization? (1mk)
5. In order to determine the molar heat of neutralization of sodium hydroxide, 100cm3 of 1M sodium hydroxide and 1M of hydrochloric acid both at the initial temperature were mixed and stirred continuously using a thermometer. The temperature of the resulting solution was recorded after every 30seconds until the highest temperature was attained. Thereafter the temperature of the solution was recorded for a further two minutes.
6. Why was it necessary to stir the mixture of the two solutions? (1mk)
7. Write an ionic equation for the reaction that took place. (1mk)
8. The sketch below was obtained when temperature of the mixture was plotted against time. Study it and answer the questions that follow.

 Y2

Y3

Temperature (oC)

 Y1

 Time (**Seconds**)

Explain the temperature changes between points

Y1 and Y2 (1mk)

Y2 and Y3 (1mk)

1. If the initial temperature for both solution was 25oC and the highest temperature was 31.4oC for the mixture. **Calculate;**

Heat change for the reaction (Specific heat capacity of solution=42Hg-1K-1, Density of the solution =1gcm-3) (2mks)

Molar heat of neutralization of sodium hydroxide. (2mks)

1. **Explain** how the molar heat of neutralization obtained in this experiment would compare with one that would be obtained using 1.0M ethanoic acid and 100cm3 of 1M sodium hydroxide solution. (2mks)

**Draw** an Energy level diagram for the reaction represented by reaction between hydrochloric acid and sodium hydroxide solution. (3mks)

1. Give the name of **one** reagent which when reacted with concentrated hydrochloric acid produces chlorine gas. (1mk)
2. The set up below was used to prepare iron (III) chloride using the apparatus shown in the diagram below.

 Guard tube

 Iron fillings

Dry chlorine gas

 Iron(III) Chloride

 Heat

1. State **one** precaution that should be taken in carrying out the above experiment. (1mk)
2. Explain why

Calcium oxide would be preferred to calcium chloride in the guard tube. (2mks)

It is necessary to pass chlorine gas through the apparatus before heating begins. (2mks)

1. Write a **chemical** equation for the reaction that took place in the guard tube.(1mk)
2. What property of Iron (III) chloride makes it possible to be collected as shown in the diagram? (1mk)
3. During the reaction in the combustion tube, the total mass of iron (III) chloride formed was found to be 1.5g. **Calculate** the volume of chlorine gas that reacted with iron. (Fe=56.0, CI=35.5 and molar gas volume at 298k is 24,000cm3) (3mks)
4. **Draw** and **name** the structure of the compound formed when excess chlorine gas is reacted with ethane gas. (2mks)
5. State **one** use of chlorine gas. (1mk)
6. Give the systematic names for the following compounds
7. HCOOCH2CH3 (1mk)
8. CH3CH2CH2CHCH2 (1mk)
9. CH C CH2 CH3 (1mk)
10. Study the flow chart below and use it to answer the questions that follow

Pro-I-ene

 CI2(g)

M

 Step I

 Sodium metal

Products

Propan-ol

 Step III

 Step II

Organic compound K

CH3CH2COONa(գ)

H2O(i)

CO2(g)

CH3CH2COOH

 Na2CO3(s)

 Mixture of Na OH and Ca(OH)2

1. Identify the organic compound **K** (1mk)
2. Write the formula of **M** (1mk)
3. Give **one** reagent that can be used in

Step I (1mk)

Step II (1mk)

1. Write the equations for the reaction in **step II** (1mk)
2. The structure below represents a type of cleansing agent.

 SO3-Na+

R

**Describe** how the cleansing agent removes grease from a piece of cloth. (3mks)

1. The set up below was used to collect gas F, produces by the reaction between water and calcium metal

 Test tube

Gas F

 Beaker

 Calcium metal

Water

1. Name gas **F** (1mk)
2. At the end of the experiment, the solution in the beaker was found to be a week base. **Explain** why the solution is a weak base. (2mks)
3. Give **one** laboratory use of the solution formed in the beaker (1mk)
4. When excess calcium metal was added to 50cm3 of 2 M aqueous copper (II)nitrate in a beaker, a brown solid and bubbles of gas were observed.
5. Write **two** equations for the reactions which occurred in the beaker. (2mks)
6. **Explain** why it is not advisable to use sodium metal for this reaction. (2mks)
7. **Calculate** the mass of calcium metal reacted with copper(II)nitrate solution (Relative atomic mass of Ca=40 (2mks0
8. Write the **formula** of the complex Ion formed in each of the reactions below.
9. Lead metal dissolves in hot alkaline solution. (1mk)
10. Zinc hydroxide dissolves ammonia solution. (1mk0
11. Give the name of each of the processes described below which takes place when the salts are exposed to air for some time.
12. Anhydrous copper (II) sulphate becomes wet. (1mk)
13. Iron (III) chloride forms an aqueous solution. (1mk)
14. Fresh crystals of sodium carbonate decahydrate become covered with a powder of solution of carbonate monohydrate. (1mkl)
15. A certain hydrate salt has the following composition by mass. Iron 20.2%, sulphur 11.5%, water 45.5% and the rest oxygen. Its relative formula mass is 278.
16. **Determine** the formula of the hydrated salt (Fe=56,S=52, O=16, H=1) (3mks)
17. 6.95g of the hydrated salt were dissolved in distilled water and the total volume made to 250cm3 of solution. **Calculate** the concentration of the salt solution. (2mks)
18. The table below shows solubility of potassium nitrate and lead nitrate

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| **Temperature oC** | **0** | **20** | **40** | **60** | **80** | **100** |
| Solubility of KNO3 in 100g of H2O | 12.5 | 32.5 | 62.5 | 110.0 | 137.5 |  |
| Solubility of Pb(NO3)in 100g of H2O | 37.5 | 52.5 | 69.0 | 87.5 | 110.0 | 131.0 |

1. Draw the solubility curves for both salts on the same axis. (Temperature on the x-axix) (3mks)

1. A solution of lead nitrate contains 90g of the salt dissolved in 100g of water at 100oC. This solution is allowed to cool to 25oC

At what temperature will crystals first appear? (1mk)

What mass of crystals will be present at 25oC (1mk)

1. Which of the **two** salts is more soluble at 30oC (1mk)
2. Determine the concentration of lead nitrate in moles per litre when the solubility of the two salts are the same. **(Pb=207.0, O=16.0, K=39.0, N=14.0)** (3mks)