

**233/3 SAMIA SUB-COUNTY 2021 CONFIDENTIAL**

***Apart from the usual laboratory fittings, each student should have the following:***

1. About 0.5g of solid A in a stoppered container
2. About 0.5g of solid B in a stoppered container
3. 100cm<sup>3</sup> of solution R
4. 150cm<sup>3</sup> of solution Q
5. 150cm<sup>3</sup> of solution P
6. Distilled water in a wash bottle
7. About a spatula end –full of solid Calcium Hydroxide
8. Red and blue litimus papers
9. Three 250ml conical flasks
10. One burette0-50ml
11. One pipette 25ml
12. One 50ml measuring cylinder
13. One 10ml measuring cylinder
14. One 25cm<sup>3</sup> volumetric flask
15. Phenolphthalein indicator
16. Labels (2)
17. Stop watch
18. Two boiling tubes
19. One metallic spatula
20. Five test tubes on a test-tube rack
21. Wooden splint
22. Test tube holder
23. 100ml glass beaker
24. 1g sodium hydrogen carbonate
25. White paper
26. Filter funnel

***The student should also get access to:***

1. 10% Hydrogen peroxide (freshly prepared + dropper).
2. 2M Barium nitrate solution +dropper
3. 0.5M Hydrochloric acid + dropper
4. Source of heat
5. Acidified potassium manganite (VII)
6. Acidified potassium dichromate (VI)
7. 2M dilute sulphuric (VI) acid
8. Ethanol with a dropper

**NOTES**

- *Solid A is Hydrated ferrous ammonium sulphate*
- *Solution B is Melleic acid*
- *Solution R is prepared by weighing exactly 4.8g of sodium carbonate dissolves it to make 1dm<sup>3</sup> of solution*
- *Solution Q is prepared by weighing exactly 172cm<sup>3</sup> of hydrochloric acid (35-37% sp.gr 1.18)and dissolving to make 1d<sup>m</sup> of solution*
- *Solution P is prepared by weighing exactly 37.2g of sodium thiosulphste pentahydrate and dissolving to make 1dm<sup>3</sup> of solution.*