

CHEMISTRY



PAPER 3

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(a) Each student should be supplied with the following

1. Burette
2. Pipette
3. Pipette filler
4. Filter funnel
5. White tile
6. Clamp and stand
7. 2 conical flask
8. **100 cm³** glass Beaker (empty)
9. Stop watch
10. **100cm³** measuring cylinder
11. **10cm³** measuring cylinder
12. **250 cm³** volumetric flask
13. Metallic spatula
14. **6** clean test tubes
15. Test tube holder
16. **500ml** distilled water
17. White piece of paper or filter paper
18. **1** filter paper
19. **1** labelling paper
20. Phenolphthalein indicator
21. About **90cm³** solution **K**
22. About **100cm³** solution **L**
23. About **70cm³** solution **N**
24. About **90cm³** solution **P**
25. About **0.5g NaHCO₃**
26. About **1.0g** solid **Q**
27. About **0.5g** solid **R**

(b) Each student should have access to the following solutions:

1. Mean of heating
2. 2M NaOH
3. 2M HNO₃
4. Pb(NO₃)₂
5. Acidified KMNO₄
6. Potassium iodide solution

NB: the above solutions should be supplied with a dropper each.

(c) SOLUTIONS PREPARATION AND SOLID MEASUREMENTS

1. SOLUTION K IS 1M H₂SO₄
2. SOLUTION L IS 0.36 M NaOH CONTAINING 14.4g OF NAOH IN 1 LITRE OF THE SOLUTION
3. SOLUTION N IS 2M HCl
4. SOLUTION P IS 0.16 M Na₂S₂O₃(*SODIUM THIOSULPHATE*)
5. SOLID Q IS A MIXTURE OF SODIUM SULPHITE(Na₂SO₃) AND LEAD (II) CARBONATE (PbCO₃) MIXED IN THE **RATIO 1:1** (*should be thoroughly mixed*)
6. SOLID R IS MALEIC ACID