CHEMISTRY PAPER
FORM 1
TERM 3 2017
2017

MARKING SCHEME

1. (a) To circle the 2 dots to the right.
   (b) Line above – solvent font
   Line below – base line
2.
   - Add water to the mixture, warm, stir
   - Filter the mixture, the residue is copper(II)oxide
   - Evaporate water and sodium chloride remains
3.
   (a)
   P – sublimation
   Q – condensation
   R – melting
   S – deposition
   (b)
   - Dry ice
   - Iodine
4.
   (a) Increase, zinc react with air to form zinc oxide.
   (b) Decrease – carbonate decompose to zinc oxide and carbon(IV)oxide which escapes into the air.
5.
   i)
   ii) Measure the temperature of the distillate coming out.
   iii) C , it has the least boiling point, hence it boils faster.
6. To draw dropper.
7.
   a) Sodium bromine
   b) Zinc and sulphur
   c) Lead and oxygen
   d) Magnesium and nitrogen
   e) Potassium and iodine
8.

Element	Symbol
Carbon	**C**
Nitrogen	N
Oxygen	**O**
Hydrogen	H
Copper	**Cu**
Magnesium	Mg

9.
- Transparent, hence reactions can be observed easily.
- Does not react with most reagents.

10.
- Does not make beakers dirty because it has no soot.
- It is the hottest hence heats faster.

11.
a) 30^0C

b)

PQ – Temperature increases steadily as solid X absorbs heat energy, the heat absorbed increases kinetic energy of the particles as they vibrate vigorously.
QR – solid-liquid Temperature remains constant until solid X melts heat supplied is used to weaken forces of attraction holding particle of solid X together. The particles move far apart as solid changes from solid to
RS – Temperature rises steadily as the liquid X absorbs heat energy. The heat supplied increases further the kinetic energy of the particle causing them to move fast.

12.
- Bhang
- Cocaine
- Miraa

13. Wash immediately with cold water.

14.
- Don't run in the lab.
- Label reagents to avoid confusion.
- Don't smell gases directly, waft instead.

15.
(a) Parts of the flame
A – pale blue zone
B – green blue zone
C – almost colourless zone

(b)