## 233/3

## CHEMISTRY PRACTICAL

## TERM TWO

CONFIDENTIAL
FORM THREE
In addition to the apparatus and fittings found in the Chemistry Laboratory, each candidate will require the following;

1. $150 \mathrm{~cm}^{3}$ of solution C
2. $150 \mathrm{~cm}^{3}$ of solution $D$
3. $60 \mathrm{~cm}^{3}$ of Copper (II) sulphate ( 2 molar)
4. $50 \mathrm{~cm}^{3}$ Burette
5. $25 \mathrm{~cm}^{3}$ pipette
6. Atleast 2 conical flasks
7. 100 ml plastic beaker
8. $\left(-10^{0} \mathrm{C}=110^{\circ} \mathrm{C}\right)$ thermometer
9. Spatula
10. One boiling tube
11. 5 test tubes
12. 1 g of Zinc powder (accurately measured)
13. About 0.5 g of solid Q
14. Tissue paper
15. Distilled water in a wash bottle

## Access to;

- Phenolphthalein indicator
- Source of heat
- 2 M sodium hydroxide with a dropper
- 2 M ammonia solution with a dropper
- 2 M hydrochloric acid with a dropper
- 2 M Barium nitrate with a dropper
- 2 M nitric (v) acid with a dropper
- 2 M Lead (II) nitrate with a dropper


## NOTES

1. Solution $C$ is made by dissolving $10.08 \mathrm{~g} / \mathrm{l}$ of ethanedioic acid (oxalic acid) in $400 \mathrm{~cm}^{3}$.
2. Solution D is made by dissolving 8.0 g of Sodium hydroxide in $400 \mathrm{~cm}^{3}$ of water and made to one litre.
3. Solid Q - aluminum sulphate $\left(\mathrm{Al}_{2}\left(\mathrm{SO}_{4}\right)_{3}\right)$
