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**CHEMISTRY**

**PAPER 3**

**PRACTICAL**

**FORM FOUR**

**ENDTERM 1 EXAM - 2021**

**CONFIDENTIAL**

In addition to the apparatus and fittings found in a Chemistry Laboratory, each candidate will require the following:

1. About 75cm3 of solution M

2. About 100cm3 of solution N

3. A burette

4. A pipette

5. Two conical flasks

6. About 1.5g of metal P

7. About 1.6g of metal Q

8. About 40cm3 of solution S

9. About 40cm3 of solution

10. A thermometer (0 - 110oC)

11. 2 plastic beakers 100ml

12. A measuring cylinder 50ml

13. About 10cm3 of solution U

14. About 0.5g of solid J

**ACCESS TO:**

1. 2M Sodium hydroxide

2. 2M Ammonium hydroxide

3. 1M Lead II Nitrate Supplied with a dropper

4. 0.5M Barium Nitrate

5. 0.5M Hydrochloric acid

**NOTES:**

1. Solution M is prepared by dissolving 3.2g of KMnO4 in 400cm3 of Sulphuric acid and diluting to

one litre of solution using distilled water.

2. Solution N is prepared by dissolving 16.7g of FeSO4.7H2O in 400cm3 of 1MH2SO4 and diluting

to one litre of solution using distilled water. The solid should be dissolved in the Sulphuric acid

immediately after weighing.

3. Metal T is magnesium powder.

4. Metal Q is Iron powder

5. Solution S is 0.5M FeSO4

6. Solution U is 0.5M AlCl3

7. Solid J is maleic acid