TEACHER.CO.KE SERIES 18

233/3
CHEMISTRY
PAPER 3
CONFIDENTIAL INSTRUCTIONS TO SCHOOLS

The information contained in this paper is to enable the head of the school and the teacher in charge of chemistry to make adequate preparations for this year's mock exams. NO ONE ELSE should have access to this paper or acquire knowledge of its contents. Great care MUST be taken to ensure that the information herein does not reach the candidates either directly or indirectly. The teacher in charge of chemistry should NOT perform any of the experiments in the same room as the candidates nor make the results of the experiments available to the candidates or give any other information related to the experiments to the candidates. Doing so will constitute an examination irregularity which is punishable.
In addition the apparatus and fittings found in a chemistry laboratory, each candidate will require the following;

- $150cm^3$ of solution A
- $100cm^3$ of solution B
- A burette
- A $25cm^3$ pipette
- A pipette filler
- A stand and a clamp
- 2 conical flasks (250 mls)
- Phenolpthalein indicator supplied with a dropper
- Filter funnel
- $500cm^3$ of distilled water in a washing bottle
- Thermometer (0^0C – 110^0C
- Exactly 2.4g of solid V
- Source of heat
- About 1g of solid E
- A spatula
- Boiling tube
- $10cm^3$ measuring cylinder
- 2 filter papers
- 6 test tubes in a test -tube rack
- About 1g of solid K
- About 0.5g of sodium carbonate
- Test – tube holder

ACCESS TO

- 2M Sodium hydroxide
- 2M Ammonia Solution
- 0.5M Lead (II) Nitrate
- 0.25M Acidified Barium Nitrate
- 1M Nitric (V) acid
- Acidified Potassium dichromate VI
- Bromine water

} Each supplied with a dropper

NOTES

1. Solid E is a mixture of 0.5g Iron (III) Chloride ($FeCl_3$) and 1g of Copper (II) Oxide (CuO)
2. Solid V is Potassium Chlorate ($KClO_3$)
3. Solid K is Butanic acid
4. Solid A is Oxalic acid
5. Solution A is prepared by dissolving 8.9g of Oxalic acid in $500cm^3$ of distilled water and diluting to one litre of solution.

6. Solution B is prepared by dissolving 4.72g of Sodium Hydroxide in 500cm^3 of distilled water and diluting to one litre of solution.
7. Acidified Barium Nitrate is prepared by dissolving 67.75g of Barium Nitrate in 200cm^3 of 2M Hydrochloric acid and diluting to one litre of solution using distilled water.