233/3 CHEMISTRY PRACTICAL
CONFIDENTIAL INTSRUCTIONS TO SCHOOLS

FORM THREE

1. *Each candidate is expected to have the following*

    (a) 4.6g (weight accurately) of solid Q
    (b) 150cm^3 of solution P
    (c) 100cm^3 of solution R
    (d) 50cm^3 burette
    (e) 25.0cm^3 pipette
    (f) Phenolphthalein indicator
    (g) labels
    (h) 100cm^3 measuring cylinder
    (i) 100cm^3 beaker
    (j) 10ml measuring cylinder
    (k) Test tube rack and 6 test tubes
    (l) About 1.0g of solid x
    (m) Solid N-1.0g

2. *Each candidate should have access to the following*
    (a) Source of heating
    (b) 2M NaOH
    (c) 2M H_2SO_4
    (d) 0.5M $Pb(NO_3)_2$
    (e) 2M $NH_{3(aq)}$
    (f) 1.0g of solid $NaHCO_3$
    (g) Red and blue litmus papers
    (h) Acidified Potassium Manganate (VII) solution
    (i) Ethanol
    (j) Conc H_2SO_4
    Distilled water in wash bottles
    Funnel

Note

1. Solid X- Maleic acid
2. Solid N-Alluminium Chloride ($AlCl_3$)
3. Solid Q –Zinc Carbonate
4. Solution P is 2M hydrochloric acid. Is prepared by dissolving $200cm^3$ of distilled topping it up to 1 litre with distilled water.
5. Solution R is 1M sodium hydroxide. It is prepared by dissolving 40g of NaOH in about $500cm^3$ of distilled water and topping it up to 1 litre with distilled water.
6. Acidified $KMnO_4$ is prepared by dissolving 3.2 g of $KMnO_4$ in water and adding $400cm^3$ of 2m H_2SO_4 then topping it to one litre with distilled water.