TEACHER.CO.KE SERIES 26

233/3 CHEMISTRY PRACTICAL
CONFIDENTIAL INTSRUCTIONS TO SCHOOLS

1.   *Each candidate is expected to have the following*

(a) 4.6g (weight accurately) of solid Q

(b) 150cm^3 of solution P

(c) 100cm^3 of solution R

(d) 50cm^3 burette

(e) 25.0cm^3 pipette

(f) Phenolphthalein indicator

(g) labels

(h) 100cm^3 measuring cylinder

(i) 100cm^3 beaker

(j) 10ml measuring cylinder

(k) Test tube rack and 6 test tubes

(l) About 1.0g of solid x

(m) Solid N-1.0g

2.   *Each candidate should have access to the following*
(a) Source of heating

(b) 2M NaOH

(c)2M H_2SO_4

(d) 0.5M $Pb(NO_3)_2$

(e) 2M $NH_{3(aq)}$

(f) 1.0g of solid $NaHCO_3$

(g) Red and blue litmus papers

(h) Acidified Potassium Manganate (VII) solution

(i) Ethanol

(j) Conc H_2SO_4

Distilled water in wash bottles

Funnel

Note

1. Solid X- Maleic acid
2. Solid N-Alluminium Chloride ($AlCl_3$)
3. Solid Q –Zinc Carbonate
4. Solution P is 2M hydrochloric acid.   Is prepared by dissolving 200cm^3 of distilled topping it up to 1 litre with distilled water.
5. Solution R is 1M sodium hydroxide. It is prepared by dissolving 40g of NaOH in about 500cm^3 of distilled water and topping it up to 1 litre with distilled water.
6. Acidified $KMnO_4$ is prepared by dissolving 3.2 g of $KMnO_4$ in water and adding 400cm^3 of 2m H_2SO_4 then topping it to one litre with distilled water.