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CHEMISTRY PRATICAL**CONFIDENTIAL INSTRUCTIONS TO SCHOOLS**

In addition to fitting and apparatus founding the chemistry laboratory, each candidate will requires the following.

- Hydrogen peroxide is prepared by measuring accurately using a clean measuring cylinder 100cm^3 of 100volume of hydrogen peroxide and add 900cm^3 of distilled water to make a litre of solution .
- Potassium iodide is prepared by weighing of accurately 8.3g and dissolve in 200cm^3 of distilled water and make the mark to 1 litre of solution.
- Sodium thiosulphate ($\text{Na}_2\text{S}_2\text{O}_3$) 0.1M is prepared by weighing accurately 15.8g and dissolve in 200cm^3 of distilled water and make the mark to 1 litre.
- Sulphuric acid is prepared by measuring 106cm^3 of concentrated sulphuric acid of density $1.08\text{g}/\text{cm}^3$ adding to 600cm^3 distilled water and make up to 1 litre of solution.
- Starch solution is prepared by dissolving 20.0g of starch powder in 100cm^3 of distilled water.