TEACHER.CO.KE SERIES 32

CHEMISTRY PRACTICAL

PAPER 3

This document must not be seen by the candidates whatsoever

CONFIDENTIAL INSTRUCTIONS TO SCHOOLS

In addition to the apparatus and fittings found in the chemistry laboratory, each candidate will require the following;

1. Exactly 3cm length of **solid A**
2. A bout 80cm^{3}of **solution B**
3. About 120cm of **solution D**
4. Means of labeling
5. One 50ml burette
6. One 25ml pipette
7. One pipette filler
8. Two 250ml conical flask
9. One 100ml clean beaker
10. One clean metallic spatula
11. One -10^{o}C to 110^{o}C thermometer
12. Six clean test tubes
13. One test tube holder
14. Two boiling tubes
15. One test tube holder
16. One 10ml measuring cylinder
17. One 250ml volumetric flask
18. A bout 1g of **solid E**
19. About 0.2g of **solid F**
20. About 1g of **solid G**
21. About 2cm^3 of calcium hydroxide solution in stopped test tube.
22. Clean glass rod
23. A bout 1g of solid sodium carbonate.
24. A bout 500cm^3 of distilled water in wash bottle.

Access to

1. Bunsen burner.
2. 2M aqueous sodium hydroxide with dropper
3. 2M aqueous ammonia supplied with dropper
4. 2M hydrochloric acid
5. Aqueous potassium dichromate (VI) supplied with dropper.
6. Methyl orange indicator supplied with dropper.
7. Phenolphthalein indicator supplied with dropper.

NOTES

1. Solution **B** is provided by adding 27cm3 of 98% concentrated sulphuric(vi) acid (sp gr 1.84) to 500cm3 of distilled water and making the solution 1 litre by adding more distilled water.
2. Solution **C** is provided by dissolving 12g of sodium hydroxide pellets in 500cm3 of distilled water and then adding more water to make 1 litre of the solution.
3. Calcium hydroxide solution is prepared by adding excess calcium hydroxide powder in cold distilled water and keeping the mixture overnight in a stopped container. The mixture is filtered in the morning of the experiment and supplied in a stopped test tube.