TEACHER.CO.KE SERIES 48

233/3

CHEMISTRY

CONFIDENTIAL

CONFIDENTIAL TO ALL SCHOOLS FOR CHEMISTRY TEACHERS

INSTRUCTION TO SCHOOLS

The information contained in this paper is to enable the head of the school and the teacher in charge of Chemistry to make adequate preparations for this year's Chemistry practical examination. **NO ONE ELSE** should have access to this paper or acquire knowledge of its contents. Great care should be taken to ensure that the information contained herein **DOES NOT** reach the candidates either directly or indirectly. The teacher in charge of Chemistry should NOT perform any of the experiment in the same room as the candidates nor make the results of the experiment available to the candidates or give any other information related to the experiment to the candidates.

Requirements for candidates.

In addition to the apparatus and fitting found in a chemistry laboratory, each candidate will require the following:

1. $100cm^3$ of solution J labelled solution J.
2. $120cm^3$ of solution K labelled solution K
3. $40cm^3$ of solution L labelled solution L.
4. $40cm^3$ of solution M labelled solution M.
5. $15cm^3$ of solution N labelled solution N.
6. One burette 50.0ml.
7. Three labels.
8. Two 200ml or 250ml empty beakers.
9. About $500cm^3$ of distilled water.
10. 10ml measuring cylinder (one)
11. One 50ml or 100ml measuring cylinder.
12. Stop watch /clock /any means of timing with second hand.
13. $10cm^3$ of solution Q.
14. About 0.5g of solid T.
15. About $5cm^3$ of solution P.

16. About 5cm^3 of solution R.
17. About 5cm^3 of solution S.
18. One boiling tube.
19. One glass rod.
20. Seven clean dry test-tubes.
21. One metallic spatula.
22. Red and blue litmus papers.
23. About 0.2g of sodium carbonate.

ACCESSS TO:-

1. Acidified potassium manganate (VII) supplied with a dropper.
2. 2M lead (II) nitrate supplied with a dropper.
3. 2M sodium Hydroxide supplied with a dropper.
4. Chlorine water.
5. Bromine water.
6. 2M Barium chloride solution.

PREPARATIONS

1. Solution J is prepared by diluting 100cm^3 of 20 volume hydrogen peroxide to one dm^3 with water.
2. Solution K is 2 M sulphuric acid.
3. Solution L is prepared by dissolving 13g of sodium thiosulphate in about 100cm^3 of distilled water and adding distilled water to make 1 litre of solution.
4. Solution M is prepared by dissolving 25g of potassium iodide in about 100cm^3 of distilled water and adding distilled water to make 1litre of solution.
5. Solution N is prepared by dissolving 1g of starch in 100cm^3 of water to make 1% of aqueous starch solution.
6. Solution Q is prepared by dissolving 30g of potassium iodide in about 500cm^3 of distilled water.
7. Solid T is oxalic acid.
8. Solution P: Dissolve 1g of Barium chloride in 10cm^3 of distilled water.
9. Solution R: Dissolve 1g of potassium iodide in 10cm^3 of distilled water.
10. Solution S: Dissolve 1g of sodium chloride in 10cm^3 of distilled water.
11. Solution Q: Same solution as in (6) above.

NOTE: - **SCHOOLS MUST COPY TABLE I AND COMPLETE IT BY CARRYING OUT THE EXPERIMENT FOR EACH SESSION TO ADVANTAGE THEIR CANDIDATES**