



TEACHER.CO.KE SERIES 51

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CHEMISTRY PRACTICAL

INSTRUCTION TO SCHOOLS

The information contained in this paper is to enable the head of the school and the teacher in charge of Chemistry to make adequate preparations for this year's Chemistry practical examination. **NO ONE ELSE** should have access to this paper or acquire knowledge of its contents. Great care should be taken to ensure that the information contained herein **DOES NOT** reach the candidates either directly or indirectly. The teacher in charge of Chemistry should **NOT** perform any of the experiment in the same room as the candidates nor make the results of the experiment available to the candidates or give any other information related to the experiment to the candidates.

Requirements. Each candidate

1. About 75cm³ of solution B.
2. About 120cm³ of solution C.
3. Solid A(3cm of magnesium ribbon)
4. Solid P(about 1g)
5. Solid E(about 1g)
6. 1 pipette(25.0cm³)
7. 2 conical flasks
8. 1 burette
9. 1 beaker 100ml
10. 1 boiling tube
11. 1 Thermometer
12. 1 spatula
13. 6 test tubes
14. Stop watch
15. 1 label

Access to

- Distilled water.
- 2M Hydrochloric acid(corked)
- 2M Ammonia solution.
- Phenolphalein indicator.
- 2M Lead (II) Nitrate.
- Universal indicator and chart.
- Sodium hydrogen carbonate (solid).
- Acidified potassium dichromate (VI).
- Bromine water.
- Source of heat.
- Absolute ethanol.

Preparations

1. Solution B: 0.7M Hydrochloric acid.
2. Solution C: 0.3M Sodium hydroxide.
3. Solid P and Solid E: