



233/3

CHEMISTRY

PAPER 3.

JULY/ AUGUST

CONFIDENTIAL

1. 180cm³ of solution N HCl in a beaker.
2. 180cm³ of solution M 0.2M NaOH in a beaker.
- 3.. 50cm³ of solution D, 2M HCl in a beaker.
4. 100cm³ of solution B, 0.1M Na₂S₂O₃ in a beaker.
5. 10cm³ of 1M HNO₃ in a boiling tube.
6. Ethanol in a stopped container.
7. 5cm³ of conc H₂SO₄ in a test tubewith a dropper.
8. Exact 1g of solid X which is F₂CO₃
9. About 1g of solid L in a dry stoppered container
10. About 1g of solid Q in a stoppered container access to:
11. Methy orange in a dropper
12. 0.5M Ca(OH)₂ in a dropper.
13. 1M Pb(NO₃)₂ in a dropper.
14. 2MI NaOH solution in a dropper.
15. Distilled water in a wash bottle.
16. Acidified KMnO₄ in a dropper.
17. About 0.5g of Na₂CO₃ per student.
18. Pipette (25cm³)
19. Burette.
20. Pipette filler.
21. 3 conical flasks.(250cm³)
22. Stand and clamp.
23. White tile.
24. 100cm³ glass beaker.
25. Thermometer (-10 to 110⁰C)
26. 10cm³ measuring cylinder.
27. 100cm³ measuring cylinder.
28. Stop watch/ clock.
29. Plain white paper.
30. 2 boiling tube.
31. 6 test tubes in a test tube rack.
32. A glass rod.
33. Metallic spatula.
34. source of heat.

N/B

1. solid X, Solid L and solid Q to be provided by the examining authority.

Preparation of solutions

Solution N is prepared by dissolving 68.8ml of concentrated Hydrochloric acid in 500ml of distilled then top up to 1litre of solution.

2. 1M HNO_3 is prepared by 66ml of conc. HNO_3 in 500ml of distilled water then top up to make 1litre of solution
3. Acidified KMnO_4 is prepared by dissolving 3.16g of KMnO_4 in 400cm^3 of 2M H_2SO_4 then topped up to 1litre of solution by distilled H_2O .
4. 2M H_2SO_4 prepared by dissolving 110cm^3 of conc. H_2SO_4 in 500ml of distilled H_2O then top to 1litre of solution.