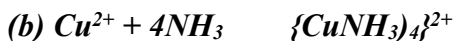


## Nitrogen and its compounds

1. (i)  $4\text{HN}_3(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$   
 (ii) Act as catalyst  
 (iii)  $\text{Zn}(\text{NH}_3)_4^{2+}$
  
2. a) Platinum/ copper  
 b) Brown fumes ✓  
 Hot rod continues to glow red  
 - NO formed reacts with oxygen to form  $\text{NO}_2$  (brown fumes)  
 - Reaction highly exothermic
  
3. a) Calcium hydroxide  
 b)  $\text{Ca}(\text{OH})_2(\text{g}) + 2\text{NH}_4\text{Cl}(\text{g}) \rightarrow 2\text{NH}_3(\text{g}) + \text{CaCl}_2 + 2\text{H}_2\text{O}(\text{L})$
  
4. (a) It neutralizes air to prevent violent combustion reaction from occurring.  
 (b) Its inert and have very low b.pt of  $-196^\circ\text{C}$   
 \*MAT
  
5. a) X is Nitrogen. ✓  
 b) It is less dense than air. ✓  
 c) – In preservation of semen in artificial insemination. ✓
  
6. a) (i) Solution A contains  $\text{Pb}^{2+}(\text{aq})$  ions ✓  
 (ii) Solution B contains  $\text{Al}^{3+}(\text{aq})$  ions. ✓  
 b) – A colourless liquid at cooler parts ✓ of test-tube is formed.  
 - A white residue remains in the test-tube. ✓
  
7. a) to expel air that is in the combustion tube so that oxygen in it does not react with hot copper ✓  
 b) brown ✓ $\frac{1}{2}$  copper metal will change to black ✓ $\frac{1}{2}$   
 c) nitrogen ✓
  
8. (a) To increase the surface area over which the reaction occurs hence increased rate of reaction. ✓  
 (b)  $\text{NH}_3$  is basic and reacts with some moles of the acid hence reduction in concentration ✓
  
9. (a) (i) The solution changes from green ✓ $\frac{1}{1}$  to brown ✓ $\frac{1}{1}$  (1 mk)  
 (ii) A brown ✓ $\frac{1}{1}$  precipitate is formed. (1 mk) 3  
 (b)  $\text{Fe}^{3+}(\text{aq}) + 3\text{OH}^-(\text{aq}) \rightarrow \text{Fe}(\text{OH})_3(\text{s})$  ✓ $\frac{1}{1}$  (1 mk)
  
10. (a) – Absorbs carbon (IV) oxide from ✓ $\frac{1}{1}$  the air. (1 mk)  
 (b)  $2\text{Cu}(\text{s}) + \text{O}_2 \rightarrow 2\text{CuO}(\text{s})$  ✓ $\frac{1}{1}$  (1 mk) 3  
 (c) Because it has the rare gases. ✓ $\frac{1}{1}$  (1 mk)
  
11. (a) Anion –  $\text{CO}_3$   
 Cation –  $\text{Cu}^{2+}$



12. (a) (i)  $\text{NH}_4\text{NO}_3 (s) \rightarrow \text{N}_2\text{O}(g) + 2\text{H}_2\text{O}(g)$   
(ii)  $\text{NH}_4\text{NO}_3$  should not be heated further if the quantity remaining is small because it may explode  
or A mixture of  $\text{NH}_4\text{Cl}$  &  $\text{KNO}_3$  can be used instead of  $\text{NH}_4\text{NO}_3$  leading to double decomposition taking place safely without explosion  
(iii) Anhydrous calcium chloride in a u-tube  
(iv) Reacts with oxygen to form brown fumes of Nitrogen (IV) Oxide  
 $2\text{N}_2\text{O}(g) + \text{O}_2(g) \rightarrow 2\text{NO}_2(g)$   
(v) – Has no colour  
- Has a slight sweet smell  
- Fairly soluble in water ✓  
- Denser than air ✓
- (b) (i) Provides a large surface area for the absorption of ammonia gas by the water or prevent “bricking” back of water  
(ii) Water would brick back into the hot preparation flask causing it to crack or break /an explosion can occur ✓  
(iii) Red litmus paper would turn to blue, blue litmus paper remains blue each ✓

13. (a) B – ammonia gas ✓1  
C - nitrogen (II) oxide (NO) ✓1  
E – water ✓1  
F – unreacted gases ✓1
- (b) The mixture of ammonia and air is passed through heated/ catalyst where ammonia (II) is oxidized to nitrogen (II) oxide. ✓1
- (c) Gases are cooled and air passed through heated/ catalyst where ammonia is further oxidized to nitrogen(IV) oxide. ✓1
- (d) Fractional distillation, ✓½  
Water with a lower boiling point ✓½ than nitric (V) acid, distills left leaving the concentrates acid.

14. a)i) Fractional distillation  
ii) Argon

- b) A Sulphur  
B Ammonia gas  
C Oteum  
D Amonium sulphate

- c) i) Finely divided iron  
ii) Vanadium (v) Oxide

d) Speeds up the rate of reaction by lowering the activation energy



f) R.M.M of  $(\text{NH}_4) = 132$

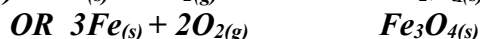
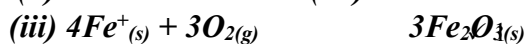
Mass of N = 28

$\% N = \frac{28}{132} \times 100 = 21.212\%$

g) Used as a fertilizer

15. (a) (i) Fused calcium chloride /Cao (quick lime)

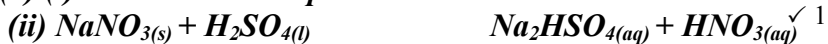
(ii) To remove carbon (IV) Oxide



(iv) Argon/Helium/Neon/Krepton

(v) Provide very low temperature so that the semen does not decompose /is not destroyed

(b) (i) Concentrated sulphuric acid



(reject unbalanced chemical equation)

(b) Copper reacts with 50% nitric acid to give nitrogen II Oxide which is colourless. Air oxidizes<sup>✓1</sup> Nitrogen II oxide to Nitrogen IV oxide which is brown.



colourless

Brown

16. (a) (i) Nitrogen – Fractional distillation of liquid air –( ½ mk)

Hydrogen – Cracking of alkanes

-Electrolysis of acidified water

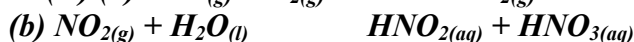
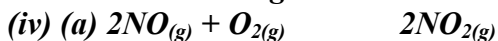
(ii) Temperature – 400°C – 500°C

Pressure – 400atm – 500atm

Catalyst – finely divided iron

(iii) Catalyst P – Nickel

Gas M – Nitrogen IV oxide



(v) To a small portion of the nitrate liquid in a test tube add equal amount of freshly prepared iron (II) sulphate followed by some drops of conc.  $H_2SO_4$  slowly on the sides. If a brown ring forms on the boundary of the two solutions, a nitrate is confirmed.

(vii) – Manufacture of nitrogenous fertilizers

- Manufacture of synthetic fibres e.g nylon

- Manufacture of explosives e.g TNT

- Manufacture of textile dyes

- Manufacture of other acids e.g. phosphoric acid

17. (a) (i) Nitrogen (I) Oxides.

Rej. Dinitrogen oxides.

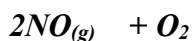


(iii) The gas is soluble in cold water.

(iv) An irritating choking smell of a gas.

(b) (i) Platinum wire.





(iii) **Nitrogen (I) Oxide**

**Colourless.**

**Relights a glowing splint.**

**Has a sweet smell.**

**Fairly soluble in water.**

**Nitrogen (IV) Oxide.**

**Reddish brown.**

**Extinguishes a glowing splint.**

**Irritating pungent smell.**

**Readily soluble in water.**

*(Accept any 1 correct comparative)*

(c) (i) **It corrodes/reacts with rubber and cork.**

(ii) **I) Oxidized : Sulphur /S**

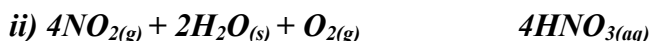
**Reduced: Nitric (V) acid / HNO<sub>(aq)</sub>**

**II) It decomposes by heat into NO<sub>2</sub> which dissolves in the acid.**

18. a) **Pass air through purifiers to remove dust particles by electrostatic precipitation. Then pass it through conc. Sodium Hydroxide to absorb CO<sub>2</sub>. Then through condensers at 25C to remove water vapour. It is further cooled to liquefy it. The liquefied air is then fractionally distilled to obtain oxygen at – 183C**

b) i) **X – Ammonia// NH<sub>3</sub>**

**Y- Air**



**Accept**



iii) **Through fractional distillation**



**RMM of NH<sub>3</sub> = 17**

**RFM of NH<sub>4</sub>NO<sub>3</sub> = 80**

**If 80g NH<sub>4</sub>NO<sub>3</sub>  $\xrightarrow{\hspace{2cm}}$  17 g**

$$\frac{960000}{80 \times 1000} \times 17 = 2040\text{kg}$$

19. (a) **Potassium hydroxide solution**

(b) **To remove dust particles**

(c) **Water vapour      Moisture**

(d) **-183°C**

(e) **Fractional distillation of liquid air**

(f) **Liquid air and passed through fractionating column, where nitrogen with lowest B.P -196°C distils out first and liquid oxygen with highest distil out last.**

(g) **Nitrogen in liquid form is used as a refrigerant e.g. in storing semen for artificial insemination**

**- Used as a raw material in Haber process e.t.c**

**II. Air is a mixture because:**

**- It contains gases which are not chemically combined**

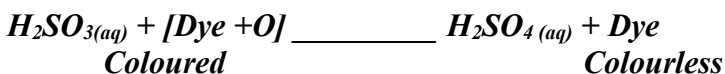
**- The gases are not in fixed ratios.**



**Coloured**

**Colourless**

**✓**



21. a) *Drying agent* ✓ ½ *which must be CaO*  
*Method of collection* ✓ - *upward delivery*  
*Workability* ✓ ½
- b)  $2\text{NH}_4\text{Cl}(\text{g}) + \text{Ca}(\text{OH})_2(\text{g}) \xrightarrow{\hspace{2cm}} \text{CaCl}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) + 2\text{NH}_3(\text{g})$  ✓
22. a) *Heat*  
 b)  $\text{Cu}(\text{g}) + \text{N}_2\text{O}(\text{g}) \xrightarrow{\hspace{2cm}} \text{CuO}(\text{g}) + \text{N}_2(\text{g})$   
 c) - *Manufacture of ammonia*  
 - *In light bulbs*  
 - *As a refrigerant*
23. - *At 113°C consists of S<sub>8</sub> rings that flow easily;*  
 - *Darkens due to breaking of S<sub>8</sub> rings and forming long chains consisting of thousands of atoms. The chains also entangle;*  
 - *The long chains consisting of thousands of atoms. The chains also entangle;*  
 - *The long chains break near b.p. to form shorter one;*
24. *Difference is at the cathode electrode where in concentrated sodium chloride sodium is deposited while in dilute sodium chloride, hydrogen is liberated, because*
25. (i)  $2\text{N}_2\text{O}(\text{g}) + \text{C}(\text{s}) \xrightarrow{\hspace{2cm}} \text{Co}_2(\text{g}) + 2\text{N}_2(\text{g})$   
 (ii) *Ammonium chloride and sodium nitrate*  
 (iii) *The hydroxide ions ✓1 (Ammonia dissolves forming ammonia hydroxide. (1 mk)*
26. (a) *E - Ammonium chloride ( ½ mk)*  
*F - Aluminium hydroxide ( ½ mk)*  
 (b)  $\text{Al}_3^+ + 3\text{OH}^-(\text{aq}) \xrightarrow{\hspace{2cm}} \text{Al}(\text{OH})_3(\text{s})$
27. a) *Zinc hydroxide*  
 b)  $[\text{Zn}(\text{NH}_3)_4]^{2+}$   
 c)  $\text{Zn}^{2+}(\text{aq}) + 2\text{OH}^-(\text{aq}) \xrightarrow{\hspace{2cm}} \text{Zn}(\text{OH})_2(\text{s})$
28. a) *Plantinum/platinum Rhodium ✓1*  
 b)  $4\text{NH}_3(\text{g}) + 5\text{O}_2(\text{g}) \xrightarrow{\hspace{2cm}} 4\text{NO}(\text{g}) \checkmark 1 + 6\text{H}_2\text{O}(\text{l})$   
 c) - *Fertilizers ✓1*  
 - *Preparation of Nitrogen (I) oxide.*  
 - *Explosives*
29. *Blue ppt ✓1 is formed which dissolves in excess to form a deep blue ✓1 solution due to formation of tetra amine Copper (II) ions*
30. (a) - *Finely divided iron impregnated by alumina (Al<sub>2</sub>O<sub>3</sub>)*  
 - *200 atmosphere pressure*  
 - *Temperature of 450°C ✓ ½*
- b) - *CuO is reduced to Copper metal*  
 - *NH<sub>3</sub> is oxidized to water and nitrogen*

31. (a) *Colour of copper (II) Oxide changes from black to brown*  
(b) (i) *Nitrogen /N<sub>2(g)</sub>*  
(ii) *Water/H<sub>2</sub>O<sub>(l)</sub>*